

Atomic Structure Review

1. What did Dalton's model look like? List the parts of his atomic theory

2. What did Rutherford discover?

3. Complete the following sentences using

sub-atomic

neutral

nucleus

conservation

matter

atoms

proton

neutron electron

a) A (n) _____ is a particle with a positive charge.

b) Most types of matter are made up of small particles called _____ .

c) _____ is anything that has mass and takes up space.

d) A (n) _____ is a particle with no charge at all.

e) A (n) _____ is a particle with a negative charge.

f) The law of _____ of matter states that matter is neither created nor destroyed.

g) Atoms are made up of _____ particles.

h) Matter that has an equal amount of positive and negative charge is said to be _____ .

i) Most of the mass of an atom is concentrated in the _____ .

Atomic Structure Review

Answers

1. What did Dalton's model look like? List the parts of his atomic theory

Dalton's model of an atom was a solid sphere



Parts of Dalton's atomic theory:

- All matter is made up of atoms that cannot be divided.
- Atoms of same elements have the same mass, atoms of different element have different masses.
- Compounds have more than one type of element
- Atoms of different elements always combine the same way in a compound

2. What did Rutherford discover?

Rutherford discovered the dense core of an atom. According to his model, protons were at the center and electrons moved around the protons on the outside.

3. Complete the following sentences using

sub-atomic
matter

neutral
atoms

nucleus
proton

conservation
neutron electron

- a) A (n) proton is a particle with a positive charge.
- b) Most types of matter are made up of small particles called atoms .
- c) Matter is anything that has mass and takes up space.
- d) A (n) neutron is a particle with no charge at all.
- e) A (n) electron is a particle with a negative charge.
- f) The law of conservation of matter states that matter is neither created nor destroyed.
- g) Atoms are made up of sub-atomic particles.
- h) Matter that has an equal amount of positive and negative charge is said to be neutral .
- i) Most of the mass of an atom is concentrated in the nucleus .